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By Erna Astuti

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Kinetic Modeling of Nitration of Glycerol

Erna Astuti^{1,2}, Supranto¹, Rochmadi¹, Agus Prasetya¹, Krister Ström³ & Bengt Andersson³

¹ Department of Chemical Engineering, Gadjah Mada University, Yogyakarta, Indonesia

² Department of Chemical Engineering, Ahmad Dahlan University, Yogyakarta, Indonesia

³ Department of Chemical and Biological Engineering, Chalmers University of Technology, Göteborg, Sweden

Correspondence: Erna Astuti, Department of Chemical Engineering, Ahmad Dahlan University, Jl. Prof. Dr. Soepomo, S. H., Janturan, Warungboto, Umbulharjo, Yogyakarta 55164, Indonesia. Tel: 62-274-563-515. E-mail: erna.astuti@che.uad.ac.id

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Abstract

Glycerol can be synthesized to make 1,3-dinitroglycerin as an intermediate product to produce polyglycidyl nitrate. This study performed kinetic modeling of nitration used elementary reactions that consisting of seven reversible reactions. The aim of this present work is to find out the representing reaction of seven reactions tested. The concept work is to study the reaction rates and its relation to the position of the nitrated hydroxyl group in the molecule. Nitration of glycerol to 1,3-dinitroglycerin was studied in the temperature range 10-30 °C, the molar ratio of nitric acid to glycerol 1/1 to 7/1 and nitric acid concentration of 69%. The seven reaction terms represent the reactions that occurred in the nitration of glycerol. The position of hydroxyl group in molecule causes difference reaction rate. The primary hydroxyl group is more reactive than secondary hydroxyl group. The parameter values describe that the 1st, 3rd, 5th and 7th reaction rates are very fast. On the contrary, the 2nd, 4th and 6th reaction rates are slow.

Keywords: nitration, glycerol, kinetic modeling

1. Introduction

Polyglycidyl nitrate (PGN) is an energetic polymer that is used as a binder in propellant (Willer & McGrath, 1997), plasticizer (Willer et al., 1995) and explosives (Braithwaite et al., 1996). PGN is the most energetic polymer (Provas, 2000). Three steps in production of PGN from glycerol are nitration, cyclization and polymerization. Glycerol is by product of biodiesel industry that need purification before was used as reactant (Ayoub & Abdullah, 2012; Kiss & Ignat, 2012). The production of PGN is an alternative process to increase economic value of biodiesel industry. This paper focuses in nitration. Nitration produces 1-MNG (1-mononitroglycerin), 2-MNG (2-mononitroglycerin), 1,3-DNG (1,3-dinitroglycerin), 1,2,4-TG (1,2-dinitroglycerin) and TNG (trinitroglycerin). It is necessary to study kinetics parameter of nitration. Basic knowledge of engineering is able to analyze and exploit fundamental reactions in production of PGN. The aim of this present work is to find out the representing reactions of nitration of glycerol. The reaction rate constants at average temperature and the activation energies will be fitted.

There have been several studies related to the nitration of glycerol. Nitration equilibrium in glycerol-aqueous system has been known (Kazakov et al., 1990a, 1990b; Yunda et al., 1991). Highsmith et al. (2002), Sanderson & Martins (2004), and Sanderson et al. (2005) proposed several examples of synthesizing glycidyl nitrate from glycerol. Then Highsmith and Johnston (2005) invented the continuous process production of glycidyl nitrate from glycerol with 1,3 dinitroglycerine as intermediate products. The kinetics parameter of nitroglycerin in the CSTR of Biazzi Process was presented by Lu et al. (2008). The temperature effect on glycerol nitration was studied by Astuti et al. (2014). Nitration of glycerol and nitric acid is a combination of both consecutive and parallel reactions (see Table 1):

Table 1. Nitration reactions

Reaction number	Reaction	ΔH_R , kJ/mole
1	$G + HNO_3 \leftrightarrow 1 \text{ MNG} + H_2O$	-59.4
2	$G + HNO_3 \leftrightarrow 2 \text{ MNG} + H_2O$	-50.6
3	$1 \text{ MNG} + HNO_3 \leftrightarrow 1,3 \text{ DNG} + H_2O$	-51.5
4	$1 \text{ MNG} + HNO_3 \leftrightarrow 1,2 \text{ DNG} + H_2O$	-46.4
5	$2 \text{ MNG} + HNO_3 \leftrightarrow 1,2 \text{ DNG} + H_2O$	-55.6
6	$1,3 \text{ DNG} + HNO_3 \leftrightarrow \text{TNG} + H_2O$	-44.4
7	$1,2 \text{ DNG} + HNO_3 \leftrightarrow \text{TNG} + H_2O$	-49.4

Source: Astuti et al., 2014.

These previous studies also described the temperature dependences of equilibrium constants of all the reactions above and the corresponding values of Gibb's energies, enthalpy, and entropy (ΔG , ΔH , ΔS) were obtained (Kazakov et al., 1990b). Rubstov and Kazakov (1997) stated that the equilibrium constants of nitration to the hydroxyl group at the secondary carbon atoms are 3 to 10 times are lower than equilibrium constants for the primary hydroxyl groups in the single-type compounds.

Order reaction is needed to arrange reaction rate equation. Nitration of some compounds have successfully been modeled as first order with respect to each reactant. The nitration of benzene and some more reactive compounds in sulphuric acid (Coombes et al., 1968), nitration of benzene, chlorobenzene, toluene, and the di- and trimethylbenzenes in trifluoroacetic acid (Moodie et al., 1977) and nitration of the same reactant in perchloric acid (Moodie et al., 1978) are the first order reaction with respect to each other. The oxidation of 2-octanol with nitric acid is also first order with respect to 2-octanol and nitric acid (Van Woezik & Westerterp, 2000). The reaction rate of trinitroglycerin formation from glycerol and nitric acid in the CSTR of Biazzini process can be expressed as (Lu et al., 2008):

$$-r_G = kC_G^n C_N^m = Ae^{-E/RT} C_G^n C_N^m \quad (1)$$

The values of n and m have been determined to be 0.9350 and 1.117 (Lu et al., 2008).

This present work describes kinetic modeling of glycerol nitration. The kinetics parameter which was calculated is rate constants for all reaction that happened in glycerol nitration at average temperature and activation energy. This paper also describes the relationship between the position of hydroxyl group and reaction rate.

2. Method

2.1 Kinetics Model Development

Referring to previous studies, in this work the obtained experimental data were treated on the hypothesis that the rate of glycerol nitration was described by first order with respect to each reactant. For many reactions involving multiple steps and pathways, the powers in the rate laws agree with the stoichiometric coefficients indicating a simple kinetic mechanism. Based on the nitration reaction (Table 1), the equation rates of nitration are defined as:

$$-\frac{dC_G}{dt} = k_1 C_G C_{HNO_3} + k_2 C_G C_{HNO_3} - k_{-1} C_{1MNG} C_{H_2O} - k_{-2} C_{2MNG} C_{H_2O} \quad (2)$$

$$-\frac{dC_{1MNG}}{dt} = k_{-1} C_{1MNG} C_{H_2O} + k_3 C_{1MNG} C_{HNO_3} + k_4 C_{1MNG} C_{HNO_3} - k_1 C_G C_{HNO_3} - k_{-3} C_{1,3DNG} C_{H_2O} - k_{-4} C_{1,2DNG} C_{H_2O} \quad (3)$$

$$-\frac{dC_{2MNG}}{dt} = k_{-2} C_{2MNG} C_{H_2O} + k_5 C_{2MNG} C_{HNO_3} - k_2 C_G C_{HNO_3} - k_{-5} C_{1,2DNG} C_{H_2O} \quad (4)$$

$$-\frac{dC_{1,3DNG}}{dt} = k_{-3} C_{1,3DNG} C_{H_2O} + k_6 C_{1,3DNG} C_{HNO_3} - k_3 C_{1MNG} C_{HNO_3} - k_{-6} C_{NG} C_{H_2O} \quad (5)$$

$$-\frac{dC_{1,2DNG}}{dt} = k_{-4} C_{1,2DNG} C_{H_2O} + k_{-5} C_{1,2DNG} C_{H_2O} + k_7 C_{1,2DNG} C_{HNO_3} - k_4 C_{1MNG} C_{HNO_3} - k_5 C_{2MNG} C_{HNO_3} \quad (6)$$

$$-\frac{dC_{NG}}{dt} = k_{-6} C_{NG} C_{H_2O} - k_6 C_{1,3DNG} C_{HNO_3} - k_7 C_{1,2DNG} C_{HNO_3} \quad (7)$$

$$\begin{aligned}
 -\frac{dC_{HNO_3}}{dt} = & k_1 C_G C_{HNO_3} + k_2 C_G C_{HNO_3} + k_3 C_{1MNG} C_{HNO_3} + k_4 C_{1MNG} C_{HNO_3} + k_5 C_{2MNG} C_{HNO_3} \\
 & + k_6 C_{1,3DNG} C_{HNO_3} + k_7 C_{1,2DNG} C_{HNO_3} - k_{-1} C_{1MNG} C_{H_2O} - k_{-2} C_{2MNG} C_{H_2O} - k_{-3} C_{1,3DNG} C_{H_2O} - \\
 & k_{-4} C_{1,2DNG} C_{H_2O} - k_{-5} C_{1,2DNG} C_{H_2O} - k_{-6} C_{NG} C_{H_2O}
 \end{aligned} \quad (8)$$

$$\begin{aligned}
 -\frac{dC_{H_2O}}{dt} = & -k_1 C_G C_{HNO_3} - k_2 C_G C_{HNO_3} - k_3 C_{1MNG} C_{HNO_3} - k_4 C_{1MNG} C_{HNO_3} - k_5 C_{2MNG} C_{HNO_3} - \\
 & k_6 C_{1,3DNG} C_{HNO_3} - k_7 C_{1,2DNG} C_{HNO_3} + k_{-1} C_{1MNG} C_{H_2O} + k_{-2} C_{2MNG} C_{H_2O} + k_{-3} C_{1,3DNG} C_{H_2O} + \\
 & k_{-4} C_{1,2DNG} C_{H_2O} + k_{-5} C_{1,2DNG} C_{H_2O} + k_{-6} C_{NG} C_{H_2O}
 \end{aligned} \quad (9)$$

The kinetics parameter can be evaluated using the experimental data. The rate constant, $k_{ref,i}$, at an average temperature (T_{ref} of 293.15 K) and the activation energy, E_A were fitted. The backward reactions for each reaction were calculated using the equilibrium constant obtained from experimental data. The Arrhenius equation for the forward and backward reactions is:

$$k_i = k_{ref,i} e^{\frac{E_A}{R} \left(\frac{1}{T_{ref}} - \frac{1}{T} \right)} \quad \text{and} \quad k_{-i} = k_i / K_i \quad (10)$$

The parameter were fitted by minimizing was the residual sums of squares of errors (SSE) for the measured concentrations of glycerol, 1-MNG, 2-MNG, 1,3-DNG, 1,2-DNG and TNG. Since steady reaction conditions were not reached immediately the first measured data point was used as a starting value when solving the differential equations describing the reactions. However, this value is also prone to measurement errors and parameters P_i was introduced to estimate the true initial condition. A least squares minimization with respect to these parameters was also introduced.

SSE is defined as:

$$SSE = \sum \frac{(Y_{calc,i} - Y_{dat,i})^2}{(MeanY_{dat,i})^2} + \sum \frac{(P_i - Y_{init,i})^2}{(MeanY_{dat,i})^2} \quad (11)$$

The weight factor $(MeanY_{dat,i})^2$ is to normalize the residual sum of squares assuming constant relative error in the concentration measurements. The function lsqnonlin with solver ode 15s in MATLAB 7.6 was used to obtain $k_{ref,i}$ and $E_{A,i}$.

2.2 Experimental Set up

The reactants used in the experiments were glycerol of $\geq 99\%$ purity from Sigma-Aldrich, nitric acid of 69% purity from Merck and 1,2-dichloroethane of purity $\geq 99.0\%$ from Sigma-Aldrich. Experiments were carried out in a 5 cm³ flask that was immersed in cooling bath. Glycerol with a certain weight was placed in flask and was diluted with an equal volume of dichloroethane and cooled to reaction temperature. Cooled nitric acid was added. The flask equipped with nitrogen purge for stirring. A thermometer monitored the temperature in the flask. Samples were taken in time intervals between 45-60 minutes, 10 samples in each experiment. Furthermore the reactions were continued until equilibrium conditions and the sample was taken. Experimental works were done with various variables i.e. reaction temperature (10 to 30 °C) and molar ratio (1/1 to 7/1).

2.3 Analytical Methods

Samples were analyzed with gas chromatography (430 GC, Bruker), equipped with flame ionization detector. The column was a VF-1 ms 30m x 0.25 mm, ID DF=1 capillary column from Factor Four. The injector and detector temperature were maintained at 175 and 225 C, respectively. The oven temperature was kept at 60 °C during injection, after that increase to 140 °C and stable at that temperature. Chromatogram were recorded by computer that used Galaxie Chromatography System version 1.9.302.952 and gave the percent mass of 1-MNG, 2-MNG, 1,3-DNG, 1,2-DNG, TNG and glycerol of each samples, respectively.

A sampling procedure was developed to minimize the time between sampling and GC analysis and to be as reproducible as possible. Still the measured concentrations show large scattering.

3. Results and Discussion

3.1 Equilibrium Constants

This paper compares the equilibrium constants which were proposed by Kazakov et al. (1990a) and equilibrium constants from experimental work. The aim of this comparison was to choose equilibrium constants that were used in simulation. The comparison can be read at Table 2.

Table 2. Equilibrium constants from Kazakov et al. (1990a) and from experimental works

Molar ratio	Temperature, °C	K ₁	K ₂	K ₃	K ₄	K ₅	K ₆	K ₇
Kazakov's	10	0.0026	1.3890	7.6074	2.0646	11.1655	0.4672	1.7210
	15	0.0026	1.2801	8.7083	2.4572	12.5396	0.4401	1.5590
	20	0.0026	1.1831	9.9226	2.9071	14.0271	0.4155	1.4171
	30	0.0026	1.0184	12.7176	4.0020	17.3587	0.3723	1.1819
1/1	10	0.0106	0.0365	6.1397	0.6715	9.4015	0.2906	2.6565
	15	0.0053	0.0304	10.1324	1.1013	2.2027	0.2448	2.2521
	20	0.0056	0.0361	11.3405	1.1674	4.6056	0.2717	2.6389
	30	0.0078	0.0329	7.4198	0.6625	1.1041	0.1512	1.6934
3/1	10	0.1351	0.1539	2.0818	0.1376	5.5055	0.4604	6.9630
	15	0.1639	0.2231	2.6187	0.2414	6.2775	0.4705	5.1034
	20	0.0760	0.3489	8.8792	0.6107	18.9329	0.0574	0.8352
	30	0.1129	0.6849	13.7496	0.9228	17.0716	0.3173	4.7273
5/1	10	0.3177	0.4347	2.6020	0.1694	15.7508	0.4692	7.2092
	15	0.4736	0.6918	2.9312	0.1937	22.6619	0.4111	6.2217
	20	0.1953	1.0903	11.4168	0.4779	1.5399	0.0839	2.0046
	30	0.5858	1.5508	5.8457	0.2682	4.1564	0.1949	4.2496
7/1	10	0.8956	1.4346	3.3170	0.1824	17.1486	0.7970	14.4904
	15	0.4690	0.8694	3.5345	0.1698	5.2632	0.3672	7.6439
	20	1.0209	1.5996	3.1964	0.1799	1.5359	0.6019	10.6957
	30	2.3883	4.0404	3.8210	0.1800	15.3039	0.5689	12.0735

Table 2 shows that Kazakov's constants do not fit properly. Equilibrium constants vary with the molar ratio. Hence, Kasakov's constants were replaced by the equilibrium constants obtained from experiments for estimation of kinetics parameter.

Experimental data have been measured in the nitration of glycerol. The nitration between glycerol and nitric acid was modeled by fitting the kinetic model with 14 parameters, $k_{ref,1}$ to $k_{ref,7}$ and $E_{A,1}$ to $E_{A,7}$ to the experimental data. Figure 1 shows the comparison between the experimental data measured and calculated data points of the six compounds that are involved in nitration. Concentrations were plotted as a function of time.

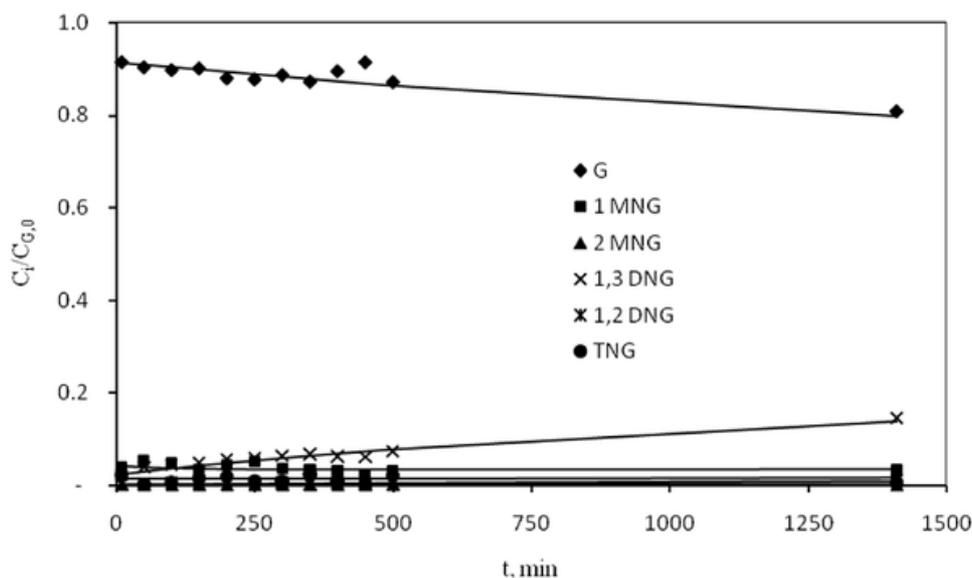


Figure 1. $C_i/C_{G,0}$ of all products for nitration of glycerol at nitric acid of 69%, molar ratio of nitric acid/glycerol 3/1 and temperature of reaction 20 °C

The nitration of glycerol produces five different compounds: two isomers of mononitroglycerin (1-MNG and 2-MNG), two isomers of dinitroglycerin (1,3-DNG and 1,2-DNG) and trinitroglycerin (TNG). Figure 1 shows that largest $C_i/C_{G,0}$ is the $C_i/C_{G,0}$ of 1,3-DNG. It is advantageous because 1,3-DNG is the main product expected from this nitration. Concentration of 2-MNG and 1,2-DNG are very small, nearest to zero. Equilibrium is reached at the beginning of the reaction. This case will be evaluated based on the estimated value of the parameter obtained from the calculation.

3.2 Influence of Temperature

Rate of reaction usually increases nearly exponentially as temperature increases. The synthesis of 1,3-DNG is an endothermic reaction (Kazakov et al., 1990b). If the reaction is endothermic (ΔH° positive) the equilibrium constant increases with temperature.

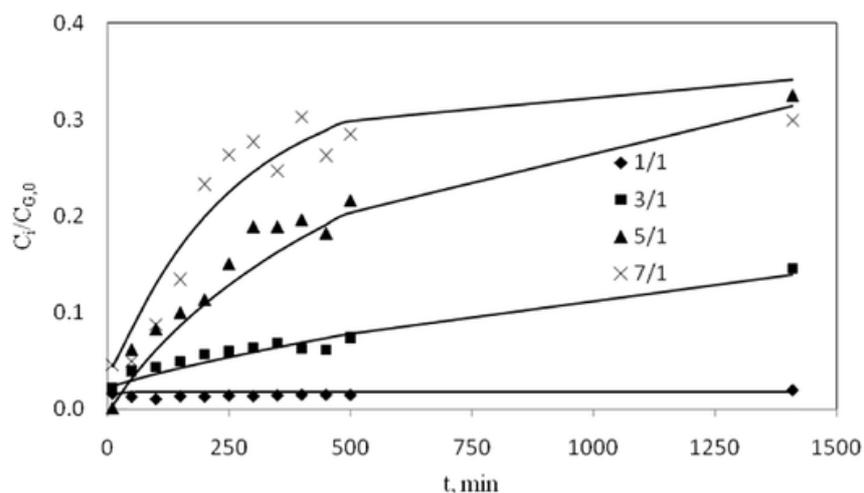


Figure 2. $C_i/C_{G,0}$ of 1,3-DNG at difference temperatures (nitric acid of 69%, molar ratio of 3/1)

It can be seen in Figure 2 that higher reaction temperature cause significant increase in the conversion of 1,3-DNG. An increase in temperature has tendency to increase the reaction rate by diminishing the resistance factor. The restriction of reaction temperature in nitration is the decomposition of trinitroglycerin that occurred at temperature above 20 °C which is dangerous for safety. The kinetic model was able to describe well the time dependence of all measured concentration especially 1,3-DNG in the low temperature (10-20 °C).

3.3 Influence of Molar Ratio

Experiments were carried out at molar ratio of nitric acid to glycerol 1/1 to 7/1. The kinetics parameter was determined separately at each molar ratio of nitric acid/glycerol. Increase in mole ratio will increase the concentration of nitric acid in the system so that the equilibrium reaction will shift to the right. Increasing mole ratio, as seen in Figure 3, clearly increased the $C_t/C_{G,0}$ of 1,3-DNG. The model is very appropriate at molar ratio 1/1 to 5/1.

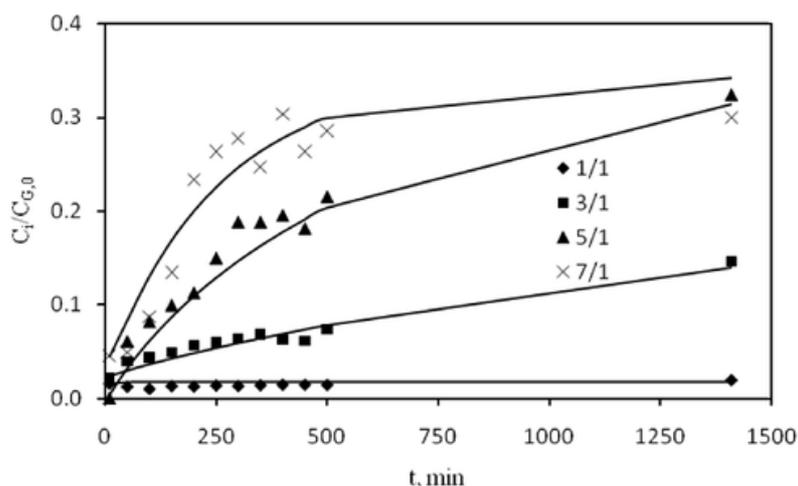


Figure 3. Experimental and estimated $C_t/C_{G,0}$ of 1,3-DNG at various mole ratio of nitric acid/glycerol (nitric acid of 69%, temperature of 20 °C)

4 is research also studied the effect of the molar ratio and the reaction temperature to yield 1,3-DNG. Increase in molar ratio of nitric acid/glycerol means increase in concentration of nitric acid. Figure 3 and Table 2 show the influence of mole ratio upon the yield of the products.

Table 3. Maximum yield of 1,3-DNG

Molar ratio	Temperature, °C	Maximum yield, %	Molar ratio	Temperature, °C	Maximum yield, %
1/1	10	1.64	5/1	10	1.57
	15	5.13		15	8.28
	20	11.05		20	22.26
	30	22.93		30	28.57
3/1	10	1.53	7/1	10	1.68
	15	6.15		15	13.42
	20	15.14		20	32.94
	30	25.09		30	38.70

As seen in Table 3, increase in molar ratio will increase the yield of 1,3-DNG. This fact supports the theory of the dependence of reaction rate on the concentration of the compounds present. At the highest molar ratio (7/1),

the deviation between the experimental yield and the yield of simulation is rather large. To some extent this is due to the high reaction rate at high concentrations of nitric acid and further reaction occurred prior to GC analysis. The same reason reliable to explain the deviation at highest reaction temperature 303.15 K.

3.4 Parameter Estimation

The resulting parameter values and their 95% confidence intervals are shown in Table 4. The sum of all nitro compounds excluding 1-MNG and 1,3-DNG was always less than 10% of all nitro compounds. Parameter estimation gave the result with arbitrary values for some reactions. Low concentration of 2-MNG and 1,2-DNG lead the accuracy were very low. As a consequence, confidence interval for the 2nd, 4th, and 6th reactions are very low. The estimation parameters (see Table 4) showed the difference between rates of nitration at the primary and secondary hydroxyl groups. The primary hydroxyl group is more reactive than secondary hydroxyl group. Reaction that fill nitrate group at first carbon atom and third carbon atom is faster than at secondary carbon atom. The 1st, 3rd, 5th and 7th reactions are the nitration of primary hydroxyl group so the reactions are fast relatively to the others. The reaction rate constants are high. Otherwise the 2nd, 4th and 6th reactions related the existence of secondary hydroxyl groups. The reaction rates are slow and the reaction rates constants are low too. This phenomenon is in line with Kazakov et al. (1990b) and Rubstov and Kazakov (1997). Kazakov et al. stated (1990b) that heat of reaction for the reaction of substitution of a primary hydroxyl is greater than for the reactions of substitution of a secondary hydroxyl. According to Rubstov et al. (1997), in polyhydric alcohols the existence of the nitrate group at the primary carbon atom increases equilibrium constants for the nitration of the secondary hydroxyl group.

It can be seen in Figure 1 the concentrations of 2-MNG and 1,2-DNG are very low. The 2nd reaction rate that including secondary hydroxyl group is very slow that causes the concentration of 2-MNG as product is very low. There are three reasons for the low concentration of 1,2-DNG. First, at the 4th reaction the active functional group is a secondary nitrate group. Therefore the formation of 1,2-DNG from 1-MNG was very slow. Second, the reaction rate of 5th reaction (the formation of 1,2-DNG from 2-MNG) is fast but the concentration of 2-MNG is very small so the concentration of 1,2-DNG as product is also very small. Third, the reaction rate of 7th reaction (the formation of TNG from 1,2-DNG) is fast. The 1,2-DNG disappear very quickly and turn to TNG.

A different experimental design with different initial compositions including some of the products would be necessary in order to obtain significant parameters for the remaining reactions. However, that was not the objective of this study. Representing reactions for the nitration reaction are known.

Table 4. Kinetic parameters for nitration of glycerol (nitric acid concentration Of 69%, 293.15K)

Reaction number	$k_{ref}, m^3/mol.s. \cdot 10^{10}$				
	1/1	3/1	5/1	7/1	all
1	1034134.55±65373133.83	5.77±0.34	15.24± 0.64	235.38±19.03	0.5±0.15
2	0.02± 0.20	0.05± 0.10	1.59±0.30	4.17±1.32	0.49±0.14
3	86632.50± 313434.38	80.18±3.22	117.22±5.72	89.97± 1.21	97.79± 7.82
4	3805.42± 8375.65	2.33± 3.65	4.90± 6.89	34.60± 8.38	26.17± 10.42
5	3.72± 156.23	59.91±110.02	10803.65±150411.43	246130.55± 13263348.16	1319.12± 1727.00
6	56.82±13.45	0.94±0.72	2.28±0.39	1.58± 0.06	30.77± 3.11
7	857.40± 282.03	2.44±7.34	103.54± 33.07	716.19± 69.27	194.64± 34.00
Activation energy, kJ/mol					
1	38.10±3692.52	38.10±4.21	38.15±2.56	38.12±5.52	38.11±13.89
2	117.23± 1441.48	117.23± 105.65	117.23± 11.62	117.23± 2.57	117.23± 12.44
3	58.62± 332.69	58.62± 3.40	58.69± 3.97	58.64± 1.20	58.62± 6.68
4	31.40± 129.65	31.40± 67.81	31.40± 79.80	31.38± 12.44	31.40± 21.67
5	167.47± 3061.09	167.48± 147.73	167.48± 406.37	167.46± 880.47	167.48± 95.45
6	104.67± 32.10	104.53± 50.04	104.51± 14.36	104.53± 2.97	104.53± 10.07
7	71.18± 31.58	71.89± 192.41	71.90± 24.04	71.84± 6.61	71.89± 14.92

The model satisfactorily captures the reactions of nitration as seen in Figures 1-3. The model gives good results at low temperature and low molar ratio. Model accuracy was not impaired by merging of the data sets obtained at four molar ratios (the result is symbolized by *all* in Table 4). The parameter values obtained for the merge data set is in the range of the parameter values which determined separately. The results obtained show the influence of composition of reactants on the reaction rate. It is better to obtain the kinetic parameter by separating of the data sets at each molar ratio.

4. Conclusion

The kinetic model adequately captures the main feature experimental work. The approach of kinetic modeling proved fruitful, indicates that all important reactions are observed. Seven reaction terms represent the reactions that occurred in the nitration of glycerol. The position of hydroxyl group in molecule causes difference reaction rates. The primary hydroxyl group is more reactive than secondary hydroxyl group. The parameter values describe that the 1st, 3rd, 5th and 7th reaction rates are very fast. On the contrary, the 2nd, 4th and 6th reaction rates are slow. All reactions represent reactions for the nitration.

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Nomenclature

A	= pre-exponential factor, $m^3/(mol)(s)$
C_i	= concentration of compound i, mol/L
E_A	= activation energy, kJ/mol
K_i	= equilibrium constant at reaction i
k_i	= rate constants of reaction i, $m^3/(mol)(s)$
$k_{ref,i}$	= rate constants of reaction i at average temperature, $m^3/(mol)(s)$
PGN	= polyglycidyl nitrat
R	= gas constants
TNG	= nitroglycerin
T	= reaction temperature, K
T_{ref}	= average temperature, K
ΔH^i	= heat of reaction, kJ/mol
1,2-DNG	= 1,2-dinitrolycerin
1,3-DNG	= 1,3-dinitrolycerin
1-MNG	= 1-mononitrolycerin
2-MNG	= 2-mononitrolycerin

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